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number. This is also the number of protons in each atom.

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In the mid-nineteenth century, chemists came to the conclusion that elements should be organized by their atomic weights. However, the atomic weights of various elements were calculated erroneously, and chemists also observed some

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anomalies in the properties of other elements. Over time, it became clear that the periodic table as currently comprised contained gaps, missing elements that had yet to be discovered. A rush to discover these missing pieces followed, and a seemingly



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Endless amount of elemental discoveries were proclaimed and brought into laboratories. It wasn't until the discovery of the atomic number in 1913 that chemists were able to begin making sense of what did and what did not belong on the

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periodic table, but even then, the discovery of radioactivity convoluted the definition of an element further. Throughout its formation, the periodic table has seen false entries, good-faith errors, retractions, and dead ends; in fact, there have been

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