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Chapter Test B, continued 17. A substance combines with oxygen, releasing a large amount of energy as heat and light, in a(n) . 18. The decomposition of a substance by an electric current is called . 19. A(n) orders the elements by the ease with which they undergo certain chemical reactions. 20.

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Chapter Test B, continued 7. The discovery of the nucleus was a result of Rutherford's observation that a small percentage of the positively charged particles bombarding the metal's surface a. were slightly deflected as they passed through the metal.

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Measurements and Calculations TEST B 1. a 2. c 3. c 4. a 5. c 6. c 7. a 8. a 9. time 10. mass 11. density 12. energy or work 13. length or distance 14. volume 15. area 16. qualitative 17. quantitative 18. qualitative 19. quantitative 20. 3.00 ± 10^{-5} km/s 21. three 22. 2.6 ± 10^{-2} g 23. 2.50 ± 10^{-2} L 24. quantity 25.

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Modern Chemistry 141 Chapter Test Name Class Date

Chapter Test B, continued 18. When a strong acid is titrated with a weak base, the pH of the solution at the equivalence point is than 7. 19. A is a highly purified solid used to check the concentration of a standard solution. 20. A 1 M solution of NaOH will have a pH that is

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Modern Chemistry 33 Chapter Test Name Class Date Chapter Test B, continued 15. The energy state of an atom is called its ground state. 16. The number of waves that pass a point in one second is called. 17. When an electron drops from a higher-energy state to a lower-energy state, a(n) spectrum is produced. 18.

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Modern Chemistry 91 Chapter Test Name Class Date Chapter Test B, continued 27. Distinguish between ionic crystals and metallic crystals. PART V Write the answers to the following questions on the line to the left, and show your work in the space provided. Modern Chemistry Chapter 12 Review Answer Key

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Assessment ...

Modern Chemistry 20 Chapter Test Name Class Date
Chapter Test A, continued ____ 13. To calculate the number of atoms present in 2.0 mol of an element, you would a. add Avogadro's number of atoms per mole to 2.0 mole. b. subtract Avogadro's number of atoms per mole from 2.0 mole. c.

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Chapter Test B, continued 16 Modern chemistry
chapter 3 test b answers. The measure of the ability of an atom in a chemical compound to attract electrons from another atom in the compound is called _____. 17. The energy required to remove one electron from an atom is called its _____. 18.

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CHAPTER 3 TEST continued Date Class FILL IN THE
BLANK Write the correct term (or terms) in the space
provided. 9, If a particular ompound is composed of
elements A and B, the ratio of the mass of B to the
mass of A will alw ys be the same. This is a statement
of the law of exactly 12 g of carbon-12 is referred to
as a(n) 11.

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(b) positive ions. (d) negative ions. 3. b Compared with the neutral atoms involved in the formation of an ionic compound, the crystal lattice that results is (a) higher in potential energy. (c) equal in potential energy. (b) lower in potential energy. (d) unstable. 4. b The lattice energy of compound A is greater in magnitude than that of ...

~~6 Chemical Bonding~~

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b. Explain how an atom can exist in this state. Atoms consist of a positively charged nucleus, made up of protons and neutrons, that is surrounded by a negatively charged electron cloud. The positive and negative charges combine to form a net neutral charge. MODERN CHEMISTRY ATOMS: THE BUILDING BLOCKS OF MATTER 19

~~3 Atoms: The Building Blocks of Matter~~

Modern Chemistry 137 Chapter Test Name Class Date Chapter Test A, continued ____19. In the figure on the previous page, the pH at the equivalence point a. is equal to 7.0. b. is greater than 7.0. c. is less than 7.0. d. cannot be determined from the data given. ____20. In the figure on the previous page, the volume of titration standard

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CHAPTER 22 TEST Nuclear Chemistry Class MULTIPLE CHOICE On the line at the left of each statement, write the letter of the choice that best completes the statement or answers the question. After converting units, the nuclear mass defect is equivalent to the a. atomic mass b. electrostatic force c. energy of chemical reaction

~~San Ramon Valley High School~~

CHAPTER 16 REVIEW Reaction Energy SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. For the following examples, state whether the change in entropy favors the forward or reverse reaction: forward reaction a. $\text{HCl(l)} \rightarrow \leftarrow \text{HCl(g)}$ reverse reaction b. $\text{C}_6\text{H}_{12}\text{O}_6(\text{aq}) \leftrightarrow \text{C}_6\text{H}_{12}\text{O}_6(\text{s})$ forward reaction c. $2\text{NH}_3(\text{g}) \dots$

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