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every year and therefore a student's ability to achieve good marks on this topic will have a big impact on their overall success. The lesson looks at the electrolysis of both molten salts and aqueous solutions and guides students through predicting the products at each of the electrodes as well as writing half equations to represent these reactions.

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Add Lead (II) Bromide into a beaker and heat so it will turn molten, allowing ions to be free to move to conduct an electric charge. Add two graphite rods as the electrodes and connect this to a power pack or battery. Turn on power pack or battery and allow electrolysis to take place. Negative bromide ions move to the positive electrode (anode) and lose two electrons to form bromine molecules.

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Once the candidates have researched and written up their own plan in Section 1 of the ISA they should carry out their investigation providing that this is workable, valid, safe, and manageable in the laboratory. Candidates need to test the hypothesis.

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